

1. How many atoms are in 3 moles of Au?

2. How many moles are in 3.7×10^{26} atoms Fe?

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$$\text{N. } \frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$$

2. How many moles are in 3.7×10^{26} atoms Fe?

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2. How many moles are in 3.7×10^{26} atoms Fe?

$$\text{B. } \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ atoms}}$$

1. How many atoms are in 3 moles of Au?

N. $\frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$

$1.8 \times 10^{24} \text{ atoms}$

2. How many moles are in 3.7×10^{26} atoms Fe?

B. $\frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ atoms}}$

1. How many atoms are in 3 moles of Au?

N. $\frac{6.02 \times 10^{23} \text{ atoms}}{1 \text{ mole}}$

$1.8 \times 10^{24} \text{ atoms}$

2. How many moles are in 3.7×10^{26} atoms Fe?

614.6 mol

B. $\frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ atoms}}$

3. How many molecules are in 2.5 moles of H_2O ?

4. How many moles are in 8.9×10^{30} molecules of CH_4 ?

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K. 6.02×10^{23} molecules
1 moles

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K. 6.02×10^{23} molecules
1 moles

4. How many moles are in 8.9×10^{30} molecules of CH_4 ?

M. 1 mole
 6.02×10^{23} molecules

3. How many molecules are in 2.5 moles of H_2O ?

K. $\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ moles}}$

$1.5 \times 10^{24} \text{ molecules}$

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$$1.5 \times 10^{24} \text{ molecules}$$

4. How many moles are in 8.9×10^{30} molecules of CH_4 ?

$$\text{M. } \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ molecules}}$$

$$1.5 \times 10^7 \text{ mol}$$

5. How many atoms are in 2.5 moles of H_2O ?

6. How many moles are in 4.5×10^{31} atoms of CH_4 ?

5. How many atoms are in 2.5 moles of H_2O ?

$$\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} \times \frac{3 \text{ atoms}}{1 \text{ molecule}}$$

6. How many moles are in 4.5×10^{31} atoms of CH_4 ?

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$$\text{C. } \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} \times \frac{3 \text{ atoms}}{1 \text{ molecule}}$$

6. How many moles are in 4.5×10^{31} atoms of CH_4 ?

$$\text{O. } \frac{1 \text{ molecule}}{5 \text{ atoms}} \times \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ molecules}}$$

5. How many atoms are in 2.5 moles of H_2O ?

$$\text{C. } \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} \times \frac{3 \text{ atoms}}{1 \text{ molecule}}$$

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$$\text{O. } \frac{1 \text{ molecule}}{5 \text{ atoms}} \times \frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ molecules}}$$

$$1.5 \times 10^7 \text{ mol}$$

7. How many grams are in 2 moles of $\text{Al}(\text{OH})_3$?

8. How many grams are in 3 moles of CaCO_3 ?

9. How many grams are in 2.3 moles of Na ?

7. How many grams are in 2 moles of $\text{Al}(\text{OH})_3$?

A. $\frac{78.0 \text{ g}}{1 \text{ mole}}$

8. How many grams are in 3 moles of CaCO_3 ?

9. How many grams are in 2.3 moles of Na ?

7. How many grams are in 2 moles of $\text{Al}(\text{OH})_3$?

A. $\frac{78.0 \text{ g}}{1 \text{ mole}}$

8. How many grams are in 3 moles of CaCO_3 ?

J. $\frac{100.1 \text{ g}}{1 \text{ mole}}$

9. How many grams are in 2.3 moles of Na?

7. How many grams are in 2 moles of $\text{Al}(\text{OH})_3$?

A. 78.0 g
1 mole

8. How many grams are in 3 moles of CaCO_3 ?

J. 100.1 g
1 mole

9. How many grams are in 2.3 moles of Na?

D. 23.0 g
1 mole

7. How many grams are in 2 moles of $\text{Al}(\text{OH})_3$?

A. $\frac{78.0 \text{ g}}{1 \text{ mole}}$

156.0 g

8. How many grams are in 3 moles of CaCO_3 ?

J. $\frac{100.1 \text{ g}}{1 \text{ mole}}$

9. How many grams are in 2.3 moles of Na?

D. $\frac{23.0 \text{ g}}{1 \text{ mole}}$

7. How many grams are in 2 moles of $\text{Al}(\text{OH})_3$?

A. $\frac{78.0 \text{ g}}{1 \text{ mole}}$

156.0 g

8. How many grams are in 3 moles of CaCO_3 ?

J. $\frac{100.1 \text{ g}}{1 \text{ mole}}$

300.3 g

9. How many grams are in 2.3 moles of Na?

D. $\frac{23.0 \text{ g}}{1 \text{ mole}}$

7. How many grams are in 2 moles of $\text{Al}(\text{OH})_3$?

A. $\frac{78.0 \text{ g}}{1 \text{ mole}}$

156.0 g

8. How many grams are in 3 moles of CaCO_3 ?

J. $\frac{100.1 \text{ g}}{1 \text{ mole}}$

300.3 g

9. How many grams are in 2.3 moles of Na?

D. $\frac{23.0 \text{ g}}{1 \text{ mole}}$

52.9 g

10. How many moles are in 789.2 g of Na?

11. How many moles are in 500.5 g of CaCO_3 ?

12. How many moles are in 400g of Al(OH)_3 ?

10. How many moles are in 789.2 g of Na?

I. 1 mole
23.0 g

11. How many moles are in 500.5 g of CaCO_3 ?

12. How many moles are in 400g of Al(OH)_3 ?

10. How many moles are in 789.2 g of Na?

I. 1 mole

23.0 g

11. How many moles are in 500.5 g of CaCO_3 ?

G. 1 mole

100.1 g

12. How many moles are in 400g of $\text{Al}(\text{OH})_3$?

10. How many moles are in 789.2 g of Na?

I. 1 mole
23.0 g

11. How many moles are in 500.5 g of CaCO_3 ?

G. 1 mole
100.1 g

12. How many moles are in 400g of $\text{Al}(\text{OH})_3$?

P. 1 mole
78.0 g

10. How many moles are in 789.2 g of Na?

I. 1 mole
23.0 g

34.3 mol

11. How many moles are in 500.5 g of CaCO_3 ?

G. 1 mole
100.1 g

12. How many moles are in 400g of $\text{Al}(\text{OH})_3$?

P. 1 mole
78.0 g

10. How many moles are in 789.2 g of Na?

I. 1 mole
23.0 g

34.3 mol

11. How many moles are in 500.5 g of CaCO_3 ?

G. 1 mole
100.1 g

5.0 mol

12. How many moles are in 400g of $\text{Al}(\text{OH})_3$?

P. 1 mole
78.0 g

10. How many moles are in 789.2 g of Na?

$$\text{I. } \frac{1 \text{ mole}}{23.0 \text{ g}}$$

34.3 mol

11. How many moles are in 500.5 g of CaCO_3 ?

$$\text{G. } \frac{1 \text{ mole}}{100.1 \text{ g}}$$

5.0 mol

12. How many moles are in 400g of $\text{Al}(\text{OH})_3$?

$$\text{P. } \frac{1 \text{ mole}}{78.0 \text{ g}}$$

5.1 mol

13. How many liters are in 4 moles of CO_2 at STP?

14. How many liters are in 3.4 moles of N_2 at STP?

15. How many moles are in 46.7 L of H_2 at STP?

13. How many liters are in 4 moles of CO_2 at STP?

E. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

14. How many liters are in 3.4 moles of N_2 at STP?

15. How many moles are in 46.7 L of H_2 at STP?

13. How many liters are in 4 moles of CO_2 at STP?

E. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

14. How many liters are in 3.4 moles of N_2 at STP?

F. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

15. How many moles are in 46.7 L of H_2 at STP?

13. How many liters are in 4 moles of CO_2 at STP?

E. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

14. How many liters are in 3.4 moles of N_2 at STP?

F. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

15. How many moles are in 46.7 L of H_2 at STP?

Q. $\frac{1 \text{ mole}}{22.4 \text{ L}}$

13. How many liters are in 4 moles of CO_2 at STP?

E. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

89.6 L

14. How many liters are in 3.4 moles of N_2 at STP?

F. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

15. How many moles are in 46.7 L of H_2 at STP?

Q. $\frac{1 \text{ mole}}{22.4 \text{ L}}$

13. How many liters are in 4 moles of CO_2 at STP?

E. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

89.6 L

14. How many liters are in 3.4 moles of N_2 at STP?

F. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

76.2 L

15. How many moles are in 46.7 L of H_2 at STP?

Q. $\frac{1 \text{ mole}}{22.4 \text{ L}}$

13. How many liters are in 4 moles of CO_2 at STP?

E. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

89.6 L

14. How many liters are in 3.4 moles of N_2 at STP?

F. $\frac{22.4 \text{ L}}{1 \text{ mole}}$

76.2 L

15. How many moles are in 46.7 L of H_2 at STP?

Q. $\frac{1 \text{ mole}}{22.4 \text{ L}}$

2.1 mol

16. How many grams are in 46.7 L of H_2 at STP?

17. How many grams are in 32.0 L of CO_2 at STP?

16. How many grams are in 46.7 L of H_2 at STP?

$$\text{L. } \frac{1 \text{ mole}}{22.4 \text{ L}} \times \frac{2 \text{ g}}{1 \text{ mole}}$$

17. How many grams are in 32.0 L of CO_2 at STP?

16. How many grams are in 46.7 L of H_2 at STP?

$$\text{L. } \frac{1 \text{ mole}}{22.4 \text{ L}} \times \frac{2 \text{ g}}{1 \text{ mole}}$$

17. How many grams are in 32.0 L of CO_2 at STP?

$$\text{H. } \frac{1 \text{ mole}}{22.4 \text{ L}} \times \frac{44.0 \text{ g}}{1 \text{ mole}}$$

16. How many grams are in 46.7 L of H_2 at STP?

$$\text{L. } \frac{1 \text{ mole}}{22.4 \text{ L}} \times \frac{2 \text{ g}}{1 \text{ mole}}$$

4.2 g

17. How many grams are in 32.0 L of CO_2 at STP?

$$\text{H. } \frac{1 \text{ mole}}{22.4 \text{ L}} \times \frac{44.0 \text{ g}}{1 \text{ mole}}$$

16. How many grams are in 46.7 L of H_2 at STP?

$$\text{L. } \frac{1 \text{ mole}}{22.4 \text{ L}} \times \frac{2 \text{ g}}{1 \text{ mole}}$$

4.2 g

17. How many grams are in 32.0 L of CO_2 at STP?

$$\text{H. } \frac{1 \text{ mole}}{22.4 \text{ L}} \times \frac{44.0 \text{ g}}{1 \text{ mole}}$$

62.9 g